



## Hankuk University of Foreign Studies

### 2019 Summer Session

### CHEM101 Introduction to Chemistry with Lab

#### Course Outline

**Term:** July 01-August 02,2019

**Class Hours:**12:00-13:50 (Monday through Friday)

**Course Code:** CHEM101

**Instructor:** Dr. Aleksandar Jeremic

**Home Institution:** The George Washington University

**Office Hours:** Tuesdays and Thursdays (15:00-16:00) or By Appointment

**Email:** jerema@gwu.edu

**Credit:** 4

**Class Hours:** This course will have 72 class hours, including 40 lecture hours, 10 lecturer office hours, 10-hour TA discussion sessions, 2-hour review sessions, 10 laboratory hours.

#### Course Description

*Introduction to Chemistry* will provide students with an overview of the current trends and body of knowledge in Chemistry, including basics of the scientific method and of the analysis of scientific data.

#### Course Objectives

The main course goal is to allow students to reach a comprehensive understanding of the issues and methods in Chemistry, in order to decide whether to pursue studies in the field. In the process of reaching this goal, our objectives are that each student will:



- Become familiar with current scientific theories and research in the major topic areas of Chemistry.
- Discover the personal relevance of course material in their everyday and professional lives, in order to make fully informed decisions.
- Develop the skills necessary to evaluate and think critically about information concerning chemical and biological phenomena obtained from research, the general public, and the media.
- Be well prepared for advanced courses in Chemistry/Life Sciences.

### **Required Textbooks**

-*Beginning Chemistry*, David Ball, (version 1.0)

Download for free at <https://2012books.lardbucket.org/pdfs/beginning-chemistry.pdf>

Publisher: LardBucket (*Creative Commons* licensed)

Publish Date: December 28, 2012

Several readings will be required throughout the course, either to prepare for class or to complete an assignment. Additional material will be posted online to provide a free and easy access to everyone.

### **Grading & Evaluation**

Lab (20%) – Midterm exam (40%) – Final exam (40%)

90-100% (A), 80-89% (B), 70-79% (C), 60-69% (D), ≤ 60% (F)

Intermediary assignments will be posted throughout the course, to help students assess their needs and to ensure that all the important topics are well understood. Assignments and labs are also an opportunity for students to ask questions concerning unclear notions, as the main objective is not to grade but to help everyone reach an optimal level of comprehension.

Midterm and final exams will target all topics previously covered in class. Lecture notes, labs and assignments are important to succeed in the midterm and final exams, yet some questions will be specifically intended to stimulate students' critical thinking. Attendance is extremely important for success in this class. It is expected that each student will commit fully to the assignments and readings required. Exams will cover the required texts as well as material presented or discussed in class.

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## Course Schedule

### **Week 1:**

Lecture 1: Course Introduction / Units and Measurements in Chemistry

Lecture 2: Atomic Theory

Lecture 3: Molecules and Chemical Nomenclature

Lecture 4: Electronic Configurations and Periodic Table

*LAB 1: Literature Review*

### **Week 2:**

Lecture 5: Masses of Atoms and Molecules

Lecture 6: Ions and Ionic Compounds

Lecture 7: Types of Chemical Reactions and Equations

Lecture 8: Stoichiometry and the Mole in Chemical Reactions

*LAB 2: Experimental design and Hypothesis testing*

### **Week 3:**

Lecture 9: Calculations and Practice Problems

Lecture 10: Review Session I

Lecture 11: **MIDTERM EXAM**

Lecture 12: Enthalpy and Chemical Reactions

*LAB 3: Collecting Data*

### **Week 4:**

Lecture 13: The Chemical Equilibrium and the Law of Mass Action

Lecture 14: Calculating Equilibrium Constant Values

Lecture 15: Chemical Bonds: Ionic Bonds

Lecture 16: Chemical Bonds: Covalent Bonds

*LAB 4: Data Analysis*

### **Week 5:**

Lecture 17: Properties of Solids / Liquids and Intermolecular Forces

Lecture 18: Solutions: Dilutions and Concentrations

Lecture 19: Acids, Bases, Buffers and pH

Lecture 20: Review Session II

### **FINAL EXAM**

*LAB 5: Writing scientific report*